

Prentice Hall Properties Of Gases Section Review Answers

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Prentice Hall Properties Of Gases

Properties of Gases Prentice-Hall Chapter 14.1 Dr. Yager Objectives Explain why gases are easier to compress than solids or liquids Describe the three factors that affect gas pressure Compressibility Compressibility is a measure of how much the volume of matter can decrease under pressure.

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The effusion rate of a gas is inversely proportional to the square root of its molecular mass. The volume of a gas increases as the pressure on that gas decreases.

Prentice Hall Chemistry Chapter 14: The Behavior of Gases ...

Gases are easily , or squeezed into a smaller volume 1. because of the between particles in a gas. The four variables 2. used to describe a gas are pressure, (P), (V), (T), 3. and number of (n). 4. You can use theory to predict and explain how gases 5. will respond to a change in conditions. Doubling the amount of 6. gas in a rigid container ...

05 CTR ch14 7/12/04 8:13 AM Page 347 THE PROPERTIES OF ...

Chapter Objectives The major goals of this chapter are to enable you to: Distinguish between speed and velocity. Use vectors to illustrate and solve velocity problems.

Chapter Objectives - Motion

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1B: The Kinetic Model The behavior of ideal gases is explained by the kinetic molecular theory of gases. Molecular motion, which leads to collisions between molecules and the container walls, explains pressure, and the large intermolecular distances in gases explain their high compressibility.

1: The Properties of Gases - Chemistry LibreTexts

SECTION 14.1 PROPERTIES OF GASES(pages 413–417) This section uses kinetic theory to explain

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the properties of gases. This section also explains how gas pressure is affected by the amount of gas, its volume, and its temperature. Compressibility(pages 413-414)

THE BEHAVIOR OF GASES 14

Atmosphere - envelope of gases that surrounds the planet; very thin layer. Composition of the Atmosphere - Made up of a mixture of atoms and molecules of different kinds Made up of nitrogen, oxygen, carbon dioxide, water vapor, and other gases, as well as particles of liquids and solids.

Prentice Hall - Weather and Climate

Chapter 14 The Behavior of Gases147 SECTION 14.1 PROPERTIES OF GASES(pages 413-417) This section uses kinetic theory to explain the properties of gases. This section also explains how gas pressure is affected by the amount of gas, its volume, and its temperature. Compressibility (pages 413-414) 1. Look at Figure 14.1 on page 413.

SECTION 14.1 PROPERTIES OF GASES(pages 413-417)

At constant volume, temperature and pressure are directly proportional. For example, if the temperature of a gas increases, then the pressure increases. Work Step by Step At constant volume, temperature and pressure are directly proportional.

Chapter 14 - The Behavior of Gases - 14.2 The Gas Laws ...

Variable shape Variable volume Definite volume Definite shape Solids Liquids Gases false true plasma gas Particles in a solid are packed close together in an orderly arrangement. The arrangement of particles in a liquid is more random. A gas takes the shape and volume of its container. 9.

Chapter 3 States of Matter Section 3.1 Solids, Liquids ...

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14.1 Properties of Gases >Compressibility Gases are easily compressed because of the space between the particles in a gas. • The distance between particles in a gas is much greater than the distance between particles in a liquid or solid. • Under pressure, gas particles are forced closer together.

chapter14 section01 srhs 2015 - Weebly

particles of an ideal gas is negligible compared to the total volume in which the gas is contained.

____ 11. The kinetic theory of gases assumes that, during a collision between two particles, kinetic energy is transferred without loss from one particle to the other. 9 8 7 6 5 4 3 2 1 THE PROPERTIES OF GASES SECTION REVIEW 12.1

12.1 The Properties of Gases Section Review

Real Gases Real gas, in contrast, has real volume and the collision of the particles is not elastic, because there are attractive forces between particles. As a result, the volume of real gas is much larger than of the ideal gas, and the pressure of real gas is lower than of ideal gas.

Gas Laws: Overview - Chemistry LibreTexts

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Prentice Hall Chemistry Chapter 5 Vocab

Hall Ch 13 : Prentice Hall Chemistry Chapter 13: States of Matter. Try it risk-free for 30 days. 1. The Kinetic Molecular Theory: Properties of Gases. What makes a gas ideal? What types of characteristics do ideal gases have? In this lesson ... 2. Pressure: Definition, Units, and Conversions. 3. The... Prentice Hall Chemistry Chapter 13: States ...

States Of Matter Prentice Hall Chapter Test

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