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the periodic table - an elements mass listed on the table equals the mass of one mole of that element where can you find the molar mass of an element? the periodic table - find the mass of each element in a molecule multiplied by the number of times that element appears in the molecular formula.

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atomic mass. the mass of an atom expressed in atomic mass units (AMU) mole. the standard unit used to measure the amount of a substance in a chemical reaction. molar mass. the mass in grams of 1 mole of a substance. avogadro's number.  $6.022 \times 10^{23}$ , it is the number of particles of atoms or molecules in 1 mole.

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Guide

## **How many moles of O<sub>2</sub> will react with 6 moles ... - study.com**

The mole is the most common unit used to express the quantity of a chemical substance. For all solids, liquids, and gases, you can convert mass to moles (or moles to mass). For gases, you should memorize the following conversion of volume to moles (or moles to volume): at 0°C and 1 atmosphere pressure (known as standard temperature and pressure or STP ), one mole of any gas occupies approximately 22.4 liters.

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atomic mass is the mass of 1 atom of a particular element and molar mass is the mass of 1 mole of an atom or molecule how many atoms in 1 mole of water?  $6.022 \times 10^{23} \times 3$

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chemistry i (tesc 141) study guide moles/ stoichiometry mole-a unit of measurement that expresses the amount of atoms, molecules or some other unit The number of items in one mole is commonly referred to Avogadro's number which equals

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Stoichiometry which generally used for the conversion of number of moles in mol of any given chemical species to the mass of that species in grams through molar mass is commonly known as mass...

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Mole: Mole is the SI unit for measuring the amount of substance.

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One mole can be defined as the amount of substance that contains Avogadro number of . atoms, molecules, ions or electrons.

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It is the number of atoms present is 12 amu of C-12. Where one mole of a substance contains an Avogadro's Number of particles.  
{eq}\rm 1\sim mole = 6.022\times 10^{\{23\}}\sim particles {/eq}

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Which has the smaller mass, 1 mole of  $\text{He}$  atoms or 4 moles of  $\text{H}$  atoms? Avogadro's Number: The quantity that relates the amount of particles to moles is known as Avogadro's Number.

## **Which has the smaller mass, 1 mole of He atoms or 4 moles ...**

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