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the equilibrium

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Equilibrium

Calculations Solve each
problem and show all

of your work. 1. At

equilibrium, a 5.0L

flask contains: 0.75

mol of PCl_5 0.50 mol of

H_2O 7.50 mol of HCl

5.00 mol of POCl_3

Calculate the K_{eq} for

the reaction: $\text{PCl}_5 (s) +$

$\text{H}_2\text{O} (g)$

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Equilibrium

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equilibrium Section
Resources. ... systems,
while dynamic
equilibrium is a
“closed” system.

Answer to . . . Figure
25 The arrows in the
diagram shown
represent opposing
physical ... Download a
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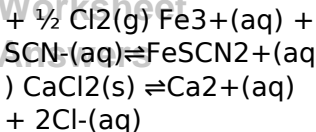
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additional

**Section 7.5 7.5
Equilibrium -
Physical Science**

Equilibrium Worksheet
SOLUTIONS. Complete
the following questions
on a separate piece of
paper. Write the
equilibrium expression,
K_{eq}, for each of the
following reactions: 2
NO(g) + O₂(g) ⇌ 2
NO₂(g) 4 HCl(g) +
O₂(g) ⇌ 2 H₂O(g) + 2
Cl₂(g) NOCl(g) ⇌ NO(g)

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equilibrium so $Q = K$.

This process is a forward and reverse reaction take place at equal rates. Such reactions are said to reach an equilibrium in which the ...

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Worksheet C

Equilibrium

Calculations . Solve

each problem and

show all of your work in

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your portfolio. 1. At equilibrium, a 2.0 L flask contains: 0.200 mol of PCl_5 0.30 mol of H_2O 0.60 mol of HCl 0.300 mol of POCl_3

Worksheet #8 Equilibrium Calculations

In an experiment, 0.75 mole of $\text{HI}(\text{g})$, 0.10 mole of $\text{H}_2(\text{g})$, and 0.50 mol of $\text{I}_2(\text{g})$ are placed in a 1.0 L container and allowed to reach equilibrium at

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1,000 K. Determine whether the equilibrium concentration of HI(g) will be greater than, equal to, or less than the initial concentration of HI(g). Justify your answer. $Q = \frac{[H_2][I_2]}{[HI]^2} =$

General Equilibrium Essay Questions FR worksheet KEY

Worksheet #9

Equilibrium

Calculations Solve each

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problem and show all of your work. 1. At equilibrium, a 5.0L flask contains: 0.75 mol of PCl_5 0.50 mol of H_2O 7.50 mol of HCl 5.00 mol of POCl_3 Calculate the K_{eq} for the reaction: $\text{PCl}_5 (s) + \text{H}_2\text{O} (g) \rightleftharpoons 2\text{HCl} (g) + \text{POCl}_3 (g)$ $K_{eq} = 23$.

2.

Worksheet #1 **Approaching** **Equilibrium**

Page 4 of 6 12. In the

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following reaction $\text{COCl}_2(\text{g}) \rightleftharpoons \text{CO}(\text{g}) + \text{Cl}_2(\text{g})$
with $K_p = 6.7 \times 10^{-9}$ at 100.0°C , if a reaction vessel is initially filled with only $\text{COCl}_2(\text{g})$ at 2.3 atm and $100.^\circ\text{C}$, what will be the equilibrium partial pressures of all three gases? 13.

115 Lab

EQUILIBRIUM

WORKSHEET Name:

a.) $2 \text{ N O}(\text{g}) \rightleftharpoons \text{N}_2(\text{g}) + \text{O}_2(\text{g})$

...

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7: Equilibria, Equilibrium

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Quiz & Worksheet - Chemical and Dynamic Equilibrium

...

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Students will be able to define chemical equilibrium and use Le Chatelier's Principle to predict how changes in pressure, concentration, and temperature affect systems in equilibrium through doing an activity, taking notes, watching videos, and performing practice questions.

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